Ber. Inst. Erdwiss. K.-F.-Univ. Graz	ISSN 1608-8166	Band 11	Graz 2006

Experimental Crystallization of $CaCO_3$ Polymorphs for Sr^{2+}/Ca^{2+} and $^{44}Ca/^{40}Ca$ Fractionation

Jianwu Tang[1], Stephan Köhler[1], Martin Dietzel[1], Albrecht Leis[2]

[1]Institute of Applied Geosciences, Graz University of Technology, Rechbauerstrasse 12, 8010 Graz, Austria

[2]Institute of Hydrogeology and Geothermics, Elisabethstrasse 16/11, 8010 Graz, Austria

Trace elements and isotopes fractionate during $CaCO_3$ crystallization. The degree of fractionation is commonly influenced by $CaCO_3$ crystallization environments and may be used as a proxy to reconstruct paleoenvironments (such as T and solution chemistry). In this study, CO_2 Diffusion Technique (CDT; e.g. Dietzel et al., 2004) and Double Diffusion Technique (DDT; e.g. Prieto et al., 1997) are used to crystallize $CaCO_3$ polymorphs. The aim is to assess the impact of temperature, precipitation rate, solution chemistry, cation diffusion, and $CaCO_3$ polymorphs on Sr^{2+}/Ca^{2+} and calcium isotope fractionation in inorganic systems.

The experimental results indicate that at a temperature range from 5º to 40ºC, single type of $CaCO_3$ polymorphs (calcite and aragonite) can be produced by controlling the aqueous Mg^{2+}/Ca^{2+} molar ratio and the precipitation rate, $R_{calcite}$ ($\mu mol\ m^{-2}\ h^{-1}$), in the reaction solution. For example, for CDT a Mg^{2+}/Ca^{2+} molar ratio less than 0.01 yields calcite as sole precipitate, whereas a ratio of about 2 exclusively induces aragonite formation. Most calcite crystals exhibit rhombohedral habit. Aragonite occurs as fibrous crystals, usually in radiating groups.

The Sr^{2+}/Ca^{2+} distribution between aqueous solution and $CaCO_3$ minerals is very sensitive to polymorphs due to the respective crystal type. Under all experimental conditions, Sr^{2+}/Ca^{2+} molar ratios in aragonite are higher than that in calcite at analogous experimental conditions. Accordingly, e.g. for calcite a continuous enrichment of aqueous Sr^{2+} vs. Ca^{2+} in reaction solution is observed, caused by Sr^{2+} discrimination in the crystal lattice, which can be followed by a Rayleigh fractionation process.

The results show that the distribution coefficient of Sr^{2+} into calcite, $D_{Sr} = (Sr^{2+}/Ca^{2+})_{calcite} / (Sr^{2+}/Ca^{2+})_{aq}$ is positively correlated to the precipitation rate ($R_{calcite}$) at a constant temperature of 5º, 25º, and 40ºC. Elevated precipitation rates usually lead to lower

discrimination effects during the precipitation (see also Lorens, 1981; Tresorio and Pankow, 1996**)**.

However, the value of D_{Sr} for calcite is also influenced by temperature. At an identical precipitation rate D_{Sr} values show a negative temperature dependence (also Dickson, 1985 and Rimstidt et al., 1998). In general, slopes for $\log(D_{Sr})$ vs. $\log(R)$ decrease from high temperature to low temperature, which indicates that precipitation rate effect is enlarged at lower temperatures. At very low precipitation rates the overall temperature effect might be small. In this case, Sr discrimination in calcite may be controlled by slow precipitation kinetics or even by equilibrium fractionation. Further experiments at low precipitation rates have to be carried out by using CDT with seed crystals to verify this behavior.

Up to now, distribution of $^{44/40}Ca$ are analyzed for several experiments at 40ºC at elevated precipitation rates. Our data lay between the results for stirred and unstirred experiments for spontaneous calcite growth at 21ºC from Lemarchand et al. (2004). Thus, preliminary results indicate that temperature may have a minor impact on $^{44/40}Ca$ isotope fractionation during calcite precipitation.

Experiments using DDT show that cation diffusion in gel follows the order of Ba>Ca>Sr>Mg and ^{40}Ca diffuses faster than ^{44}Ca at 25°C. The degree of fractionation may be caused by individual diffusion coefficients and the structure of the gel. In terms of the overall Ca isotope fractionation by calcite crystallization, diffusion effect is highly significant. Measured overall Ca isotope fractionation for calcite, $\Delta^{44/40}Ca_{calcite\text{-}Ca2+}$, from DDT is for example about -1 ‰, whereas $\Delta^{44/40}Ca_{Ca2+\ diffusion}$ due to calcium diffusion without $CaCO_3$ formation ranges from -0.4‰ to -1.5‰, depending on the diffusion length and concentration gradient.

.

References

Dickson, J.A.D., 1985. Sedimentology – Recent Developments and Applied Aspects, 173-188.

Dietzel, M., Gussone, N., Eisenhauer, A., 2004. Co-precipitation of Sr^{2+} and Ba^{2+} with aragonite by membrane diffusion of CO_2 between 10 and 50ºC. Chemical Geology 203, 139-151.

Lemarchand, D., Wasserburg, G.J., Papanastassion, D.A., 2004. Rate-controlled calcium isotope fractionation in synthetic calcite. Geochim. Cosmochim Acta 68, 4665-4678.

Lorens, R.B., 1981. Sr, Cd, Mn and Co distribution coefficients in calcite as a function of calcite precipitation rate. Geochim. Cosmochim. Acta 45, 553-561.

Prieto, M., Fernández-González, A., Putnis, A., Fernández-Díaz, L., 1997. Nucleation, growth, and zoning phenomena in crystallizing $(Ba,Sr)CO_3$, $Ba(SO_4,CrO_4)$, $(Ba,Sr)SO_4$, and $(Cd,Ca)CO_3$ solid solutions from aqueous solutions. Geochim. Cosmochim. Acta 61, 3383-3397.

Rimstidt, J.D., Balog, A., Webb, J., 1998. Distribution of trace elements between carbonate minerals and aqueous solutions. Geochim. Cosmochim. Acta 62, 1851-1863.

Tesoriero, A., Pankow, J.F., 1996. Solid solution partitioning of Sr^{2+}, Ba^{2+}, and Cd^{2+} to calcite. Geochim. Cosmochim. Acta 60, 1053-1063.